

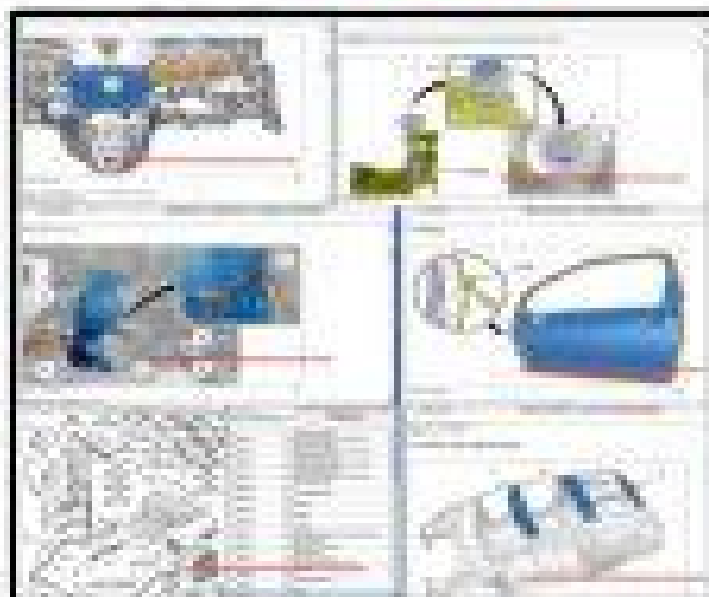
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 $\text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = 84.7 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 1560 \text{ kJ}$

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 $\text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \quad \Delta h = -1300 \text{ kJ}$

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