



CIVIL DRAFTING TECHNOLOGY

Fourth Edition

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Terence M. Shumaker



ISBN 0-02-300070-0
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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq} + 2 \text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six Na^+ ions and the six Cl^- ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$

2 of the spectator ions in KOH(aq) and $\text{HNO}_3\text{(aq)}$ are $\text{K}^+\text{(aq)}$ and $\text{NO}_3^-\text{(aq)}$ how to identify and list the

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web the $\text{Ag}^+\text{(aq)}$ and $\text{Cl}^-\text{(aq)}$ ions become AgCl(s) but the $\text{Na}^+\text{(aq)}$ ions and the $\text{NO}_3^-\text{(aq)}$ ions stay as $\text{Na}^+\text{(aq)}$ ions and $\text{NO}_3^-\text{(aq)}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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