

Chemistry Unit 3

Study Guide

Chemical Nomenclature

Chemical formula

- a way of presenting information about the chemical proportions of atoms that constitute a particular chemical compound or molecule, using chemical element symbols, numbers, and sometimes also other symbols

Subscript

- A subscript or superscript is a character that is set slightly below or above the normal line of type, respectively. It is usually smaller than the rest of the text. Subscripts appear at or below the baseline, while superscripts are above.

Ion

- Ion, any atom or group of atoms that bears one or more positive or negative electrical charges.

Anion

- a negatively charged ion

Cation

- A positively charged ion

Polyatomic ion

- The prefix poly- means many, so a polyatomic ion is an ion that contains more than one atom.

Understand the use of Roman Numerals in ionic compounds

- Roman numerals are used in naming ionic compounds when the metal cation forms more than one ion.

List the diatomic molecules

- Hydrogen (H_2)
- Nitrogen (N_2)
- Oxygen (O_2)
- Fluorine (F_2)
- Chlorine (Cl_2)
- Iodine (I_2)
- Bromine (Br_2)

Be able to name and write chemical formulas for ionic, covalent, and acidic compounds

- Properly use prefixes
- Properly use Roman Numerals
- Properly use endings (-ide, -ate, -ite, -ous, -ic)

Determine the type of compound made from a pair of elements or polyatomic ions

Identify the ionic charge of transition metals in compounds

- Because most transition metals have two valence electrons, the charge of $2+$ is a very common one for their ions.

Identify compounds as ionic, covalent, or acidic

- Compounds containing two elements (so called binary compounds) can either have ionic or covalent bonding. If a compound is made from a metal and a nonmetal, its bonding will be ionic. If a compound is made from two non-metals, its bonding will be covalent.

Chemistry Unit 3 Study Guide Answer

Arshad Iqbal



Chemistry Unit 3 Study Guide Answer:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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