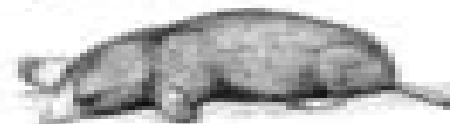


Intro to Stoichiometry – Moles to Moles NOTES

Reminder: To switch between moles of one substance and moles of another, use the ratio of the substances, which is found by looking at their coefficients.



Ex. 1: Given 3 moles of N_2 , how many moles of NH_3 would be produced?

$$\frac{3 \text{ moles } \text{N}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2}$$

Answer: 6.0 moles of NH_3



Ex. 2: Given 0.5 moles of H_2 , how many moles of NH_3 would be produced?

$$\frac{0.5 \text{ moles } \text{H}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{3 \text{ moles } \text{H}_2}$$

Answer: 0.33 moles of NH_3



Ex. 3: Given 6 moles of NH_3 , how many moles of N_2 were needed?

$$\frac{6 \text{ moles } \text{NH}_3}{1} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3}$$

Answer: 3.0 moles of N_2

Ex. 4: Given 1.5 moles of N_2 , how many moles of H_2 were needed?

$$\frac{1.5 \text{ moles } \text{N}_2}{1} \times \frac{3 \text{ moles } \text{H}_2}{1 \text{ mole } \text{N}_2}$$

Answer: 4.5 moles of H_2

Chemlab 12 A Mole Ratio

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A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the BarCharts, Inc. text.

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