

Name _____ Date _____ Class _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
 - a single atom
 - a molecule containing three atoms**
 - a molecule containing two atoms
 - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - four**
- Bond length is the distance between
 - two molecules of the same substance
 - the nuclei of two attached atoms**
 - the electrons in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - As the character of the bond, the stronger the bond.**
 - As the character of the bond, the weaker the bond.
 - As the character of the bond, the fewer the electrons in it.
 - As the character of the bond, the lower the bond dissociation energy.

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- false** Binary molecular compounds are generally composed of a metal and a nonmetal.
- true** The second element in the formula of a binary compound is named using the suffix *-ite*.
- false** The prefix *tri-* indicates three atoms.
- true** The prefix *kilo-* indicates six atoms.
- true** In naming the first element in a formula, the prefix *mono-* is not used.
- true** For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- false** An crystal contains only two elements.
- false** If the name of the cation of an ionic compound is only the acid name contains the suffix *-ate*.

In your textbook, read about naming molecular compounds and oxoacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
5. Cl_2	a. hydrobromic acid
6. CO	b. nitrogen trioxide
7. H_2CO_3	c. carbon monoxide
8. H_2SO_3	d. nitrous acid
9. NO_2	e. acetone
10. N_2O_5	f. nitric acid
11. HNO_3	g. carbonic acid
12. HNO_2	h. boric acid
13. HBr	i. carbon dioxide
14. HBrO_3	

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