

Finishing Part A
Use the chart below to determine the pH of each solution. Each pH indicator changes color based on the pH of the solution. Look at Methyl Red as an example. Methyl red will be red at and below a pH of 4.8. Additionally, methyl red will be yellow at or above a pH of 6.0. In between a pH of 4.8 and 6.0, methyl red will be orange (red and yellow make orange).

Each pH indicator has a limited range. Use the combination of the indicators to narrow down the pH range for each solution. For example, consider a solution with unknown pH. Start with the indicators that change color around a pH of 7.0, which will be bromothymol blue and phenol red. Our bromothymol blue indicator is blue, so the pH must be greater than 7.6. Likewise, the phenol red indicator is red, so the pH must be greater than 8.0. However, our phenolphthalein solution is colorless, indicating a pH lower than 8.2. Therefore, we have determined that the solution must have a pH between 8.0 and 8.2.

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- | | |
|-------------------------------------|------|
| Distilled Water | 7.0 |
| Tap Water | 6.2 |
| 0.1M NaOH | 11 |
| 0.1M NH ₄ OH | 10.1 |
| 0.1M HAc | 4.8 |
| 0.1M HCl | 4.8 |
| 0.1M H ₂ SO ₄ | 4.0 |

- | | |
|------------------------------|-------|
| Distilled Water | _____ |
| Tap Water | _____ |
| 0.1M NaOH | _____ |
| 0.1M NH_4OH | _____ |
| 0.1M HAc | _____ |
| 0.1M HCl | _____ |
| 0.1M H_2SO_4 | _____ |

Chem 1212 Lab Manual

J Spring



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Table of Contents Chem 1212 Lab Manual

1. Understanding the eBook Chem 1212 Lab Manual
 - The Rise of Digital Reading Chem 1212 Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 1212 Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 1212 Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 1212 Lab Manual
 - Personalized Recommendations
 - Chem 1212 Lab Manual User Reviews and Ratings
 - Chem 1212 Lab Manual and Bestseller Lists
5. Accessing Chem 1212 Lab Manual Free and Paid eBooks
 - Chem 1212 Lab Manual Public Domain eBooks
 - Chem 1212 Lab Manual eBook Subscription Services
 - Chem 1212 Lab Manual Budget-Friendly Options
6. Navigating Chem 1212 Lab Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem 1212 Lab Manual Compatibility with Devices
 - Chem 1212 Lab Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 1212 Lab Manual
 - Highlighting and Note-Taking Chem 1212 Lab Manual
 - Interactive Elements Chem 1212 Lab Manual
8. Staying Engaged with Chem 1212 Lab Manual

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chem 1212 Lab Manual
- 9. Balancing eBooks and Physical Books Chem 1212 Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 1212 Lab Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chem 1212 Lab Manual
 - Setting Reading Goals Chem 1212 Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem 1212 Lab Manual
 - Fact-Checking eBook Content of Chem 1212 Lab Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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