

Stoichiometric Characterization (Week of September 14th) Post-lab
Submission Guide

Directly edit this document with responses below each question. Type your answers in **BLUE** font. All answers should be typed, and all chemical reactions, calculations, and mathematical expressions should be entered using an embedded function such as Equation Editor, LaTeX, etc. When your answers are complete, upload this document to Canvas under "Assignments." *Post-labs are due by the start of your next lab period.*

Post-Lab Questions
Activity A: Determination of the Formula of a Hydrate

In the simulation in part A, you recorded the mass lost as water when $\text{CuSO}_4 \cdot x\text{H}_2\text{O}(s)$ was heated. The objective of the experiment was to use mass data to determine the number of waters of hydration. To do this, keep in mind that chemical formulas represent both atom and *mole* ratios.

Specifically, $x = \frac{\text{moles } \text{H}_2\text{O}}{\text{moles anhydrous salt}}$

Question 1

Use the data collected in activity A along with the molar masses of water and the anhydrous CuSO_4 to determine x for the hydrate. Show your calculations

$$.0657 \text{ g of } \text{H}_2\text{O} \cdot \frac{1 \text{ mol of } \text{H}_2\text{O}}{18 \text{ g of } \text{H}_2\text{O}} = .00365 \text{ mol of } \text{H}_2\text{O}$$

$$.1165 \text{ g of } \text{CuSO}_4 \cdot \frac{1 \text{ mol of } \text{CuSO}_4}{159.609 \text{ g of } \text{CuSO}_4} = .000730 \text{ mol of } \text{CuSO}_4$$

$$x = \frac{.00365 \text{ mol}}{.000730 \text{ mol}}$$

$$x = 5$$

Question 2

Determine the mass percent of water in the hydrate.

$$\text{mass percent} = \frac{.0657 \text{ g (Mass Lost)}}{.1822 \text{ g (Initial Mass)}}$$

$$\text{mass percent of water} = 36.06\%$$

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Joacim Rocklöv



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