

CHEM 102 – Recitation Week 2

Using the "Key Concepts" and the "Key Equations and Relationships" in the Chapter Summary at the end of each chapter, identify and type in this Word document some **concepts and/or equations** needed to solve each item. Be prepared to share your ideas in class discussions. You are NOT required to solve the problems in detail. If you choose to do it to better participate in class discussion, **do not forget to also identify at least 1-2 concepts to get full credit** (see the Answer Key posted for week 1, which offers examples of concept identification).

Ch. 11, 36. What type of intermolecular forces are present in each element or compound below?

- a) Kr
- b) NCl_3
- c) SiH_4
- d) HF

Concepts: • intermolecular forces
• Polarity
• Electro negativity

Ch. 11, 42. Pick the compound with the highest boiling point in each pair. Explain your reasoning.

- a) NH_3 or CH_4
- b) CS_2 or CO_2
- c) CO_2 or NO_2

Concepts: • mass of compounds
• intermolecular forces (as to which hydrogen bond must be considered)

Ch. 11, 44. For each pair given below, pick the compound with the higher vapor pressure at a give temperature. Explain your reasoning.

- a) CH_4 or CH_3Cl
- b) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ or CH_3OH
- c) CH_3OH or H_2CO

Concepts: • smaller molecules = higher BP
• Normal vapor pressure = 1 atm -

Ch. 11, 46. Determine if each pair of compounds forms a homogeneous solution (that is, are soluble in each other or miscible) when combined. For those that form homogeneous solutions, indicate the types of forces that are involved.

- a) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ and $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$
- b) CBr_4 and H_2O
- c) LiNO_3 and H_2O
- d) CH_3OH and $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

Concepts: • Intermolecular forces
• Polarity
• Tendency of dissolving

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Hugh Chisholm



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