

Section 8.4: Calculations Involving Acidic Solutions

Tutorial 1 Practice, page 513

1. Given: $[\text{HCl}(\text{aq})] = 0.0700 \text{ mol/L}$

Required: $[\text{OH}^-(\text{aq})]$

Analysis: $[\text{H}^+(\text{aq})] = 0.0700 \text{ mol/L}$ because HCl is a strong acid

$$[\text{OH}^-(\text{aq})] = \frac{1.0 \times 10^{-14}}{[\text{H}^+]}$$

$$\text{Solution: } [\text{OH}^-(\text{aq})] = \frac{1.00 \times 10^{-14}}{7.00 \times 10^{-2}}$$

$$[\text{OH}^-(\text{aq})] = 1.43 \times 10^{-13} \text{ mol/L}$$

2. Given: $[\text{HBr}(\text{aq})] = 0.0380 \text{ mol/L}$; $\text{pH} = 2.78$

Required: pH , pOH

Analysis: $[\text{H}^+(\text{aq})] = 0.0350 \text{ mol/L}$ because HBr is a strong acid

$$\text{pH} = -\log[\text{H}^+(\text{aq})]$$

$$\text{pOH} = 14.00 - \text{pH}$$

$$\text{Solution: } \text{pH} = -\log[\text{H}^+(\text{aq})]$$
$$= -\log(3.5 \times 10^{-2}) = 1.46$$

$$\text{pH} = 1.46$$

$$\text{pOH} = 14.00 - 1.46$$

$$= 12.54$$

Tutorial 2 Practice, page 516

1. Given: $[\text{HC}_2\text{H}_3\text{O}_2(\text{aq})] = 0.050 \text{ mol/L}$; $\text{pH} = 2.78$

Required: percentage ionization

Analysis: $[\text{H}^+(\text{aq})] = 10^{-\text{pH}}$

$$\frac{\text{percentage ionization}}{100} = \frac{[\text{H}^+(\text{aq})]}{[\text{HC}_2\text{H}_3\text{O}_2(\text{aq})]}$$

$$\text{Solution: } [\text{H}^+(\text{aq})] = 10^{-2.78}$$

$$[\text{H}^+(\text{aq})] = 0.00166 \text{ mol/L}$$

$$\frac{\text{percentage ionization}}{100} = \frac{0.00166}{0.050}$$

$$\text{percentage ionization} = 3.3 \%$$

2. Given: $[\text{HF}(\text{aq})] = 0.100 \text{ mol/L}$; percentage ionization = 7.8 %

Required: K_a

Analysis: $\text{HF}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{F}^-(\text{aq})$

$$K_a = \frac{[\text{H}^+(\text{aq})][\text{F}^-(\text{aq})]}{[\text{HF}(\text{aq})]}$$

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1. Understanding the eBook Chemistry12 Nelson Solution Manual
 - The Rise of Digital Reading Chemistry12 Nelson Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry12 Nelson Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry12 Nelson Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry12 Nelson Solution Manual
 - Personalized Recommendations
 - Chemistry12 Nelson Solution Manual User Reviews and Ratings
 - Chemistry12 Nelson Solution Manual and Bestseller Lists
5. Accessing Chemistry12 Nelson Solution Manual Free and Paid eBooks
 - Chemistry12 Nelson Solution Manual Public Domain eBooks
 - Chemistry12 Nelson Solution Manual eBook Subscription Services
 - Chemistry12 Nelson Solution Manual Budget-Friendly Options
6. Navigating Chemistry12 Nelson Solution Manual eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry12 Nelson Solution Manual Compatibility with Devices
 - Chemistry12 Nelson Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry12 Nelson Solution Manual
 - Highlighting and Note-Taking Chemistry12 Nelson Solution Manual
 - Interactive Elements Chemistry12 Nelson Solution Manual
 8. Staying Engaged with Chemistry12 Nelson Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry12 Nelson Solution Manual
 9. Balancing eBooks and Physical Books Chemistry12 Nelson Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry12 Nelson Solution Manual
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry12 Nelson Solution Manual
 - Setting Reading Goals Chemistry12 Nelson Solution Manual
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry12 Nelson Solution Manual
 - Fact-Checking eBook Content of Chemistry12 Nelson Solution Manual
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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